

PATENT APPLICATION

OF

ANNE MAE GAFFNEY

AND

RUOZHI SONG

FOR

**HYDROTHERMALLY SYNTHESIZED MO-V-M-NB-X OXIDE CATALYSTS FOR
THE SELECTIVE OXIDATION OF HYDROCARBONS**

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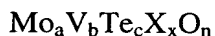
**HYDROTHERMALLY SYNTHESIZED MO-V-M-NB-X OXIDE CATALYSTS
FOR THE SELECTIVE OXIDATION OF HYDROCARBONS**

The present invention relates to a catalyst for the oxidation of alkanes, or a mixture of alkanes and alkenes, to their corresponding unsaturated carboxylic acids by vapor phase catalytic oxidation; to a method for making the catalyst; and to a process for the vapor phase catalytic oxidation of alkanes, or a mixture of alkanes and alkenes, to their corresponding unsaturated carboxylic acids using the catalyst. The present invention also relates to a process for the vapor phase catalytic oxidation of alkanes, or a mixture of alkanes and alkenes, in the presence of ammonia, to their corresponding unsaturated nitriles using the catalyst.

Nitriles, such as acrylonitrile and methacrylonitrile, have been industrially produced as important intermediates for the preparation of fibers, synthetic resins, synthetic rubbers, and the like. The most popular method for producing such nitriles is to subject an olefin such as propene or isobutene to a gas phase catalytic reaction with ammonia and oxygen in the presence of a catalyst at a high temperature. Known catalysts for conducting this reaction include a Mo-Bi-P-O catalyst, a V-Sb-O catalyst, an Sb-U-V-Ni-O catalyst, a Sb-Sn-O catalyst, a V-Sb-W-P-O catalyst and a catalyst obtained by mechanically mixing a V-Sb-W-O oxide and a Bi-Ce-Mo-W-O oxide. However, in view of the price difference between propane and propene or between isobutane and isobutene, attention has been drawn to the development of a method for producing acrylonitrile or methacrylonitrile by an ammoxidation reaction wherein a lower alkane, such as propane or isobutane, is used as a starting material, and it is catalytically reacted with ammonia and oxygen in a gaseous phase in the presence of a catalyst.

In particular, U.S. Patent No. 5,281,745 discloses a method for producing an unsaturated nitrile comprising subjecting an alkane and ammonia in the gaseous state to catalytic oxidation in the presence of a catalyst which satisfies the conditions:

(1) the mixed metal oxide catalyst is represented by the empirical formula



wherein X is at least one element selected from the group consisting of niobium, tantalum, tungsten, titanium, aluminum, zirconium, chromium, manganese, iron, ruthenium, cobalt, rhodium, nickel, palladium, platinum, antimony, bismuth, boron

and cerium and, when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $x = 0.01$ to 1.0 and n is a number such that the total valency of the metal elements is satisfied; and

(2) the catalyst has X-ray diffraction peaks at the following angles ($\pm 0.3^\circ$) of 2θ in its X-ray diffraction pattern: 22.1° , 28.2° , 36.2° , 45.2° and 50.0° .

5 Unsaturated carboxylic acids such as acrylic acid and methacrylic acid are industrially important as starting materials for various synthetic resins, coating materials and plasticizers. Commercially, the current process for acrylic acid manufacture involves a two-step catalytic oxidation reaction starting with a propene feed. In the first stage, propene is converted to acrolein over a modified bismuth
10 molybdate catalyst. In the second stage, acrolein product from the first stage is converted to acrylic acid using a catalyst composed of mainly molybdenum and vanadium oxides. In most cases, the catalyst formulations are proprietary to the catalyst supplier, but, the technology is well established. Moreover, there is an incentive to develop a single step process to prepare the unsaturated acid from its
15 corresponding alkene. Therefore, the prior art describes cases where complex metal oxide catalysts are utilized for the preparation of unsaturated acid from a corresponding alkene in a single step.

 Japanese Laid-Open Patent Application Publication No. 07-053448 discloses the manufacture of acrylic acid by the gas-phase catalytic oxidation of propene in the
20 presence of mixed metal oxides containing Mo, V, Te, O and X wherein X is at least one of Nb, Ta, W, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pd, Pt, Sb, Bi, B, In, Li, Na, K, Rb, Cs and Ce.

 Commercial incentives also exist for producing acrylic acid using a lower cost propane feed. Therefore, the prior art describes cases wherein a mixed metal oxide
25 catalyst is used to convert propane to acrylic acid in one step.

 U.S. Patent No. 5,380,933 discloses a method for producing an unsaturated carboxylic acid comprising subjecting an alkane to a vapor phase catalytic oxidation reaction in the presence of a catalyst containing a mixed metal oxide comprising, as essential components, Mo, V, Te, O and X, wherein X is at least one element selected
30 from the group consisting of niobium, tantalum, tungsten, titanium, aluminum, zirconium, chromium, manganese, iron, ruthenium, cobalt, rhodium, nickel, palladium, platinum, antimony, bismuth, boron, indium and cerium; and wherein the

proportions of the respective essential components, based on the total amount of the essential components, exclusive of oxygen, satisfy the following relationships:

$0.25 < r(\text{Mo}) < 0.98$, $0.003 < r(\text{V}) < 0.5$, $0.003 < r(\text{Te}) < 0.5$ and $0.003 < r(\text{X}) < 0.5$, wherein $r(\text{Mo})$, $r(\text{V})$, $r(\text{Te})$ and $r(\text{X})$ are the molar fractions of Mo, V, Te and X,

5 respectively, based on the total amount of the essential components exclusive of oxygen.

Nonetheless, the prior art continues to seek ways to improve the performance of such mixed metal oxide catalysts.

For example, Ueda, et al., "Hydrothermal Synthesis of Mo-V-M-O Complex
10 Metal Oxide Catalysts Active For Partial Oxidation Of Ethane", Chem. Commun., 1999, pp. 517-518, discloses Mo-V-M-O ($M = \text{Al, Fe, Cr or Ti}$) complex metal oxide catalysts prepared by hydrothermal synthesis, which exhibit activity for the partial oxidation of ethane to ethene and acetic acid.

Ueda, et al., "Selective Oxidation Of Light Alkanes Over Mo-Based Oxide
15 Catalysts", Res, Chem. Intermed., Vol. 26, No. 2, pp. 137-144 (2000) discloses that Anderson-type heteropoly compounds linked with vanadyl cations, VO^{+2} , were synthesized by hydrothermal reaction and showed good catalytic activity for the oxidation of ethene to acetic acid.

Watanabe, et al., "New Synthesis Route For Mo-V-Nb-Te Mixed Metal
20 Oxides For Propane Ammoxidation", Applied Catalysis A: General, 194-195, pp. 479-485 (2000) discloses an examination of several methods for preparing Mo-V-Nb-Te mixed oxides. Hydrothermal treatment was shown to give a precursor of an ammoxidation catalyst which shows activity twice as high, after calcination, as the catalyst prepared by the known dry-up method. Mixed oxides prepared by a solid state
25 reaction gave poor catalytic activity.

Ueda, et al., "Selective Oxidation of Light Alkanes Over Hydrothermally
Synthesized Mo-V-M-O ($M = \text{Al, Ga, Bi, Sb and Te}$) Oxide Catalysts", Applied
Catalysis A: General, 200, pp. 135-143 (2000) discloses selective oxidations of ethane
to ethene and acetic acid, and of propane to acrylic acid, carried out over
30 hydrothermally synthesized Mo-V-M-O ($M = \text{Al, Ga, Bi, Sb and Te}$) complex metal oxide catalysts. The solids were classified into two groups, i.e. Mo-V-M-O ($M = \text{Al, possibly Ga and Bi}$) and Mo-V-M-O ($M = \text{Sb and Te}$). The former were moderately active for the partial oxidation of ethane to ethene and acetic acid. The latter were

extremely active for the oxidative dehydrogenation and were also active for the partial oxidation of propane to acrylic acid. Grinding of the catalysts, after heat treatment at 600°C in N₂, increased the conversion of propane and enhanced the selectivity to acrylic acid.

5 Chen, et al., "Selective Oxidation Of Ethane Over Hydrothermally Synthesized Mo-V-Al-Ti Oxide Catalyst", *Catalysis Today*, 64, pp. 121-128 (2001) discloses synthesis of a monophasic material, Mo₆V₂Al₁O_x, by the hydrothermal method, which shows activity for gas-phase oxidation of ethane to ethene and acetic acid. Addition of titanium to the Mo₆V₂Al₁O_x oxide catalyst resulted in a marked increase of the
10 activity for the ethane selective oxidation.

It has now been found that certain novel mixed metal oxide catalysts, which may be prepared by a hydrothermal synthesis technique, may be utilized for the partial oxidation of an alkane, or a mixture of an alkane and an alkene, to produce an unsaturated carboxylic acid; or for the ammoxidation of an alkane, or a mixture of an
15 alkane and an alkene, to produce an unsaturated nitrile.

Thus, in a first aspect, the present invention provides a catalyst comprising a mixed metal oxide having the empirical formula



wherein

20 M is an element selected from the group consisting of Te and Sb;
X is an element selected from the group consisting of Li, Na, K, Rb, Cs, Mg, Ca, Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ta, Cr, W, Mn, Re, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, B, Ga, In, Pb, P, As, Sb, Bi, Se, F, Cl, Br, I, Pr, Nd, Sm and Tb, with the proviso that, when M is Sb, X cannot be Sb;

25 a, b, c, d, e and f are the relative atomic amounts of the elements Mo, V, M, Nb, X and O, respectively; and

when a = 1, b = 0.01 to 1.0, c = 0.01 to 1.0, d = 0.01 to 1.0 and 0 < e ≤ 1.0 and f is dependent on the oxidation state of the other elements;

said catalyst having been produced by the process comprising:

30 (i) admixing compounds of elements Mo, V, M, Nb and X, as needed, and a solvent comprising water to form a first admixture containing at least 2 but less than all of said elements Mo, V, M, Nb and X;

- (ii) heating said first admixture at a temperature of from 25°C to 200°C for from 5 minutes to 48 hours;
- (iii) then, admixing compounds of elements Mo, V, M, Nb and X, as needed, with said first admixture to form a second admixture containing elements Mo, V, M, Nb and X, in the respective relative atomic proportions a, b, c, d and e, wherein, when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $d = 0.01$ to 1.0 and $0 < e \leq 1.0$;
- (iv) heating said second admixture at a temperature of from 50°C to 300°C for from 1 hour to several weeks, in a closed vessel under pressure;
- (v) recovering insoluble material from said closed vessel to obtain a catalyst.

In a second aspect, the present invention provides a process for producing a catalyst comprising a mixed metal oxide having the empirical formula



wherein

M is an element selected from the group consisting of Te and Sb;

X is an element selected from the group consisting of Li, Na, K, Rb, Cs, Mg, Ca, Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ta, Cr, W, Mn, Re, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, B, Ga, In, Pb, P, As, Sb, Bi, Se, F, Cl, Br, I, Pr, Nd, Sm and Tb, with the proviso that, when M is Sb, X cannot be Sb;

a, b, c, d, e and f are the relative atomic amounts of the elements Mo, V, M, Nb, X and O, respectively; and

when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $d = 0.01$ to 1.0 , $0 < e \leq 1.0$ and f is dependent on the oxidation state of the other elements;

the process comprising:

- (i) admixing compounds of elements Mo, V, M, Nb and X, as needed, and a solvent comprising water to form a first admixture containing at least 2 but less than all of said elements Mo, V, M, Nb and X;
- (ii) heating said first admixture at a temperature of from 25°C to 200°C for from 5 minutes to 48 hours;
- (iii) then, admixing compounds of elements Mo, V, M, Nb and X, as needed, with said first admixture to form a second admixture containing elements

Mo, V, M, Nb and X, in the respective relative atomic proportions a, b, c, d, and e wherein, when a = 1, b = 0.01 to 1.0, c = 0.01 to 1.0, d = 0.01 to 1.0 and

$0 < e \leq 1.0$;

5 (iv) heating said second admixture at a temperature of from 50°C to 300°C for from 1 hour to several weeks, in a closed vessel under pressure;

(v) recovering insoluble material from said closed vessel to obtain a catalyst.

In a third aspect, the present invention provides a catalyst comprising a mixed metal oxide having the empirical formula



wherein

M is an element selected from the group consisting of Te and Sb;

X is an element selected from the group consisting of Li, Na, K, Rb, Cs, Mg, Ca,

15 Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ta, Cr, W, Mn, Re, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, B, Ga, In, Pb, P, As, Sb, Bi, Se, F, Cl, Br, I, Pr, Nd, Sm and Tb, with the proviso that, when M is Sb, X cannot be Sb;

a, b, c, d, e and f are the relative atomic amounts of the elements Mo, V, M, Nb, X and O, respectively; and

20 when a = 1, b = 0.01 to 1.0, c = 0.01 to 1.0, d = 0.01 to 1.0, $0 < e \leq 1.0$ and f is dependent on the oxidation state of the other elements;

said mixed metal oxide having an x-ray diffraction pattern with peaks at the following diffraction angles (2θ): 22.1 ± 0.3 , 26.2 ± 0.3 , 27.3 ± 0.3 , 29.9 ± 0.3 , 45.2 ± 0.3 , 48.6 ± 0.3 , and no peaks at the diffraction angles (2θ): 28.2 ± 0.3 , 36.2 ± 0.3 , 50.0 ± 0.3 .

25

In a fourth aspect, the present invention provides a process for producing a catalyst comprising a mixed metal oxide having the empirical formula



wherein

30 M is an element selected from the group consisting of Te and Sb;

X is an element selected from the group consisting of Li, Na, K, Rb, Cs, Mg, Ca,

Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ta, Cr, W, Mn, Re, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, B, Ga, In, Pb, P, As, Sb, Bi, Se, F, Cl, Br, I, Pr, Nd, Sm and Tb, with the proviso that, when M is Sb, X cannot be Sb;

a, b, c, d, e and f are the relative atomic amounts of the elements Mo, V, M, Nb, X and O, respectively; and

when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $d = 0.01$ to 1.0 , $0 < e \leq 1.0$ and f is dependent on the oxidation state of the other elements;

said mixed metal oxide having an x-ray diffraction pattern with peaks at the following diffraction angles (2θ): 22.1 ± 0.3 , 26.2 ± 0.3 , 27.3 ± 0.3 , 29.9 ± 0.3 , 45.2 ± 0.3 , 48.6 ± 0.3 , and no peaks at the diffraction angles (2θ): 28.2 ± 0.3 , 36.2 ± 0.3 , 50.0 ± 0.3 , the process comprising:

- (i) admixing compounds of elements Mo, V, M, Nb and X, as needed, and a solvent comprising water to form a first admixture containing at least 2 but less than all of said elements Mo, V, M, Nb and X;
- (ii) heating said first admixture at a temperature of from 25°C to 200°C for from 5 minutes to 48 hours;
- (iii) then, admixing compounds of elements Mo, V, M, Nb and X, as needed, with said first admixture to form a second admixture containing elements Mo, V, M, Nb and X, in the respective relative atomic proportions a, b, c, d and e, wherein, when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $d = 0.01$ to 1.0 and $0 < e \leq 1.0$;
- (iv) heating said second admixture at a temperature of from 50°C to 300°C for from 1 hour to several weeks, in a closed vessel under pressure;
- (v) recovering insoluble material from said closed vessel;
- (vi) calcining said recovered insoluble material;
- (vii) contacting said calcined recovered insoluble material with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acid and hydrogen peroxide to form a contact mixture;
- (viii) recovering insoluble material from said contact mixture to obtain a catalyst.

In a fifth aspect, the present invention provides a process for producing an unsaturated carboxylic acid which comprises subjecting an alkane, or a mixture of an alkane and an alkene, to a vapor phase catalytic oxidation reaction in the presence of a catalyst according to the first aspect of the invention.

5 In a sixth aspect, the present invention provides a process for producing an unsaturated nitrile which comprises subjecting an alkane, or a mixture of an alkane and an alkene, and ammonia to a vapor phase catalytic oxidation reaction in the presence of a catalyst according to the first aspect of the invention.

10 In a seventh aspect, the present invention provides a process for producing an unsaturated carboxylic acid which comprises subjecting an alkane, or a mixture of an alkane and an alkene, to a vapor phase catalytic oxidation reaction in the presence of a catalyst according to the third aspect of the invention.

15 In an eighth aspect, the present invention provides a process for producing an unsaturated nitrile which comprises subjecting an alkane, or a mixture of an alkane and an alkene, and ammonia to a vapor phase catalytic oxidation reaction in the presence of a catalyst according to the third aspect of the invention.

Figure 1 of the drawing is an x-ray diffraction pattern of the catalyst produced in Comparative Example 2.

20 Figure 2 of the drawing is an x-ray diffraction pattern of the catalyst produced in Comparative Example 3.

Figure 3 of the drawing is an x-ray diffraction pattern of the catalyst produced in Example 1.

Figure 4 of the drawing is an x-ray diffraction pattern of the catalyst produced in Example 2.

25 The catalyst of the first aspect of the present invention comprises a mixed metal oxide having the empirical formula



wherein

M is an element selected from the group consisting of Te and Sb;

30 X is an element selected from the group consisting of Li, Na, K, Rb, Cs, Mg, Ca, Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ta, Cr, W, Mn, Re, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, B, Ga, In, Pb, P, As, Sb, Bi, Se, F, Cl, Br, I, Pr, Nd, Sm and Tb, with the proviso that, when M is Sb, X cannot be Sb;

a, b, c, d, e and f are the relative atomic amounts of the elements Mo, V, M, Nb, X and O, respectively; and

when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $d = 0.01$ to 1 , $0 < e \leq 1.0$ and f is dependent on the oxidation state of the other elements;

5 the catalyst having been produced by the process comprising:

(i) admixing compounds of elements Mo, V, M, Nb and X, as needed, and a solvent comprising water to form a first admixture containing at least 2 but less than all of said elements Mo, V, M, Nb and X;

10 (ii) heating said first admixture at a temperature of from 25°C to 200°C for from 5 minutes to 48 hours;

(iii) then, admixing compounds of elements Mo, V, M, Nb and X, as needed, with said first admixture to form a second admixture containing elements Mo, V, M, Nb and X, in the respective relative atomic proportions a, b, c, d and e, wherein, when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $d = 0.01$ to 1.0 and $0 < e \leq 1.0$;

15 (iv) heating said second admixture at a temperature of from 50°C to 300°C for from 1 hour to several weeks, in a closed vessel under pressure;

(v) recovering insoluble material from said closed vessel to obtain a catalyst.

20 It is preferred that M is Te.

Additionally, it is preferred that X is an element selected from the group consisting of Li, K, Mg, Ca, Sc, Y, La, Cr, Mn, Re, Fe, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, Ga, In, Pb, Bi, Se, F, Cl, Br, I, Pr, Nd, Sm and Tb; more preferably X is an element selected from the group consisting of Li, K, Mg, Ca, La, Cr, Mn, Fe, Co, Rh, Ni, Pd, Pt, Zn, In, Pb, Cl, Br, I and Sm.

25 Moreover, it is preferable that, when $a = 1$, $b = 0.1$ to 0.6 , $c = 0.1$ to 0.5 , $d = 0.01$ to 0.5 and $e = 0.0001$ to 0.5 ; more preferably that, when $a = 1$, $b = 0.3$ to 0.6 , $c = 0.15$ to 0.5 , $d = 0.05$ to 0.25 and $e = 0.001$ to 0.25 . The value of f, i.e. the amount of oxygen present, is dependent on the oxidation state of the other elements in the catalyst.

30

In the second aspect of the present invention, the catalyst is formed by a “hydrothermal” technique. In this “hydrothermal” technique:

compounds of elements Mo, V, M, Nb and X, as needed, and a solvent comprising water are mixed to form a first admixture containing at least 2 but less than all of said elements Mo, V, M, Nb and X;

the first admixture is heated at a temperature of from 25°C to 200°C for from 5 minutes to 48 hours;

then, compounds of elements Mo, V, M, Nb and X, as needed, are mixed with the first admixture to form a second admixture containing elements Mo, V, M, Nb and X, in the respective atomic proportions a, b, c, d, and e, wherein, when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $d = 0.01$ to 1.0 and $0 < e \leq 1.0$;

the second admixture is heated at a temperature of from 50°C to 300°C for from

1 hour to several weeks, in a closed vessel under pressure; and

insoluble material is recovered from the closed vessel to obtain a catalyst.

Optionally, albeit preferably, the recovered insoluble material is calcined.

After calcination, the calcined insoluble material recovered from the closed vessel has an x-ray diffraction pattern (as measured using Cu - K α radiation as the source) with peaks at the following diffraction angles (2θ): 22.1 ± 0.3 , 27.3 ± 0.3 , 28.2 ± 0.3 , 36.2 ± 0.3 , 45.2 ± 0.3 , 50.0 ± 0.3 . The intensity of the x-ray diffraction peaks may vary depending upon the measuring conditions of each crystal. However, as long as the above-noted pattern is observable, the basic crystal structure is the same even if other peaks are observed in addition to those specified as defining the above-noted pattern.

In particular, compounds of elements Mo, V, M, Nb and X and a solvent comprising water are mixed to form a first admixture. This first admixture contains at least 2, but less than all, of the elements Mo, V, M, Nb and X needed to make the desired catalyst composition, wherein M is an element selected from the group consisting of Te and Sb, and X is an element selected from the group consisting of Li, Na, K, Rb, Cs, Mg, Ca, Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ta, r, W, Mn, e, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, B, Ga, In, Pb, P, As, Sb, Bi, Se, F, Cl, Br, I, Pr, Nd, Sm and Tb.

For the catalyst containing the elements Mo, V, M, Nb and X, it is preferred to mix compounds of the elements Mo, M, Nb and X, compounds of the elements Mo

and M or compounds of the elements V and M to form the first admixture. It is preferable to utilize the components of the first admixture in the atomic proportions of the elements present in the final catalyst.

The so-formed first admixture may be held at a temperature of 25°C to 200°C, preferably 50°C to 175°C, most preferably 75°C to 150°C, for from 5 minutes to 48 hours, preferably 5 minutes to 10 hours, most preferably 5 minutes to 5 hours. Any solvent or any portion of the solvent that evaporates from the admixture during this treatment may be condensed and returned to the admixture, if so desired, in order to maintain a liquid phase. Alternatively, this treatment of the first admixture may be carried out in a closed container.

Subsequent to the aforementioned treatment of the first admixture, the first admixture is mixed with compounds of the elements Mo, V, M, Nb and X, as needed, to form a second admixture containing elements Mo, V, M, Nb and X, in the respective atomic proportions a, b, c, d, and e wherein, when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $d = 0.01$ to 1.0 and $0 < e \leq 1.0$. Preferably, only compounds of elements missing from the first admixture are added in this mixing procedure.

The second admixture is then held at a temperature of 50°C to 300°C, preferably 100°C to 250°C, most preferably 150°C to 200°C, for from 1 hour to several weeks, preferably 2 hours to 7 days, most preferably 5 to 72 hours, preferably in a closed vessel. The closed vessel may be under pressure. The pressure may be the autogenous pressure of the enclosed materials at the particular temperature utilized or it can be an applied pressure, such as could be obtained by feeding a pressurized gaseous atmosphere to the closed vessel. In either case, a gas space over the enclosed admixture may comprise an oxidizing atmosphere such as air, oxygen enriched air or oxygen; a reducing atmosphere such as hydrogen; an inert atmosphere such as nitrogen, argon, helium or mixtures thereof; or mixtures thereof. Moreover, the gas space over the enclosed admixture may be charged with low levels of catalyst promoters such as NO_x compounds, SO_x compounds, gaseous halides or halogens.

After the aforementioned treatment, insoluble material is recovered from the closed vessel. This insoluble material is a very active catalyst, but the selectivity for the desired reaction product, e.g., acrylic acid, is generally poor. When the recovered insoluble material is subjected to calcination, the activity is decreased but the

selectivity for the desired reaction product, e.g., acrylic acid, is increased. The activity of the catalyst can be restored to some extent, while maintaining the enhanced selectivity, by contacting the calcined catalyst with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide.

Such contacting may be effected without any particular restrictions. The liquid contact member is normally used in an amount of 1 to 100 times the volume of the calcined recovered insoluble material, preferably 3 to 50 times the volume, more preferably 5 to 25 times the volume. Contacting may be effected at elevated temperatures, however, if prolonged contact times are not a consideration, contacting at lower temperatures may be utilized, e.g., contacting at temperatures less than 150°C. Normally, contact temperatures of room temperature to 100°C are utilized, preferably 50°C to 90°C, more preferably 60°C to 80°C. As previously noted, contact time will be affected by the temperature at which the contacting is carried out.

Normally, contact times of 1 to 100 hours are utilized, preferably 2 to 20 hours, more preferably 5 to 10 hours. The contact mixture is preferably agitated during the contacting.

There are no particular restrictions upon the organic acids which may be used as the liquid contacting member. For example, oxalic acid, formic acid, acetic acid, citric acid and tartaric acid may be used, however, oxalic acid is preferred. If the organic acid is a liquid, it may be used as is or in an aqueous solution. If the organic acid is a solid, it is used in an aqueous solution. When using aqueous solutions, there are no particular restrictions on the concentration of the organic acid. Normally, the concentration of the organic acid in the aqueous solution can vary from 0.1 to 50% by weight, preferably 1 to 15% by weight.

There are no particular restrictions upon the alcohols which may be used as the liquid contacting member. For example, methanol, ethanol, propanol, butanol, hexanol and diols may be utilized, however, alcohols having one to four carbon atoms are preferred, with ethylene glycol being particularly preferred. The alcohols may be utilized in the form of aqueous solutions, but, if so, the water content should be held to 20% by weight or less for the best effectiveness.

Similarly, there are no particular restrictions upon the inorganic acids which may be used as the liquid contacting member. For example, nitric acid, sulfuric acid,

phosphoric acid, hydrochloric acid, perchloric acid, chloric acid and hypochlorous acid may be used. The inorganic acids are typically used as aqueous solutions with concentrations of the acids in the range of from 0.1 to 50% by weight, preferably from 0.1 to 10% by weight.

5 When hydrogen peroxide is used as the liquid contacting member, it is used in the form of an aqueous solution having a concentration in the range from 0.1 to 50% by weight, preferably from 1 to 10% by weight.

 After contacting with the liquid contacting member, insoluble material is recovered from the so-formed contact mixture. The insoluble material may be
10 recovered by any conventional method, e.g., centrifugation or filtration. If the contacting was conducted at elevated temperature, the contact mixture may be cooled prior to recovery of the insoluble material.

 After contacting with the liquid contacting member, the insoluble material recovered from the contact mixture has an x-ray diffraction pattern (as measured using
15 Cu - K α radiation as the source) with peaks at the following diffraction angles (2θ): 22.1 ± 0.3 , 26.2 ± 0.3 , 27.3 ± 0.3 , 29.9 ± 0.3 , 45.2 ± 0.3 , 48.6 ± 0.3 , and no peaks at the diffraction angles (2θ): 28.2 ± 0.3 , 36.2 ± 0.3 , 50.0 ± 0.3 . The intensity of the x-ray diffraction peaks may vary depending upon the measuring conditions of each crystal. However, as long as the above-noted pattern is observable, the basic crystal
20 structure is the same even if other peaks are observed in addition to those specified as defining the above-noted pattern.

 In the "hydrothermal" technique, water is preferably used as the solvent in the catalyst preparation. The water is any water suitable for use in chemical syntheses including, without limitation, distilled water and de-ionized water. However, the
25 solvent utilized may further comprise polar solvents such as, for example, water-miscible alcohols, diols, glycols, polyols, ethers, carboxylates, nitriles and amides. The amount of solvent utilized is not critical.

 Preferably, the vessel utilized for hydrothermal treatment is formed of a material inert to the hydrothermal reaction or is lined with a material inert to the
30 hydrothermal reaction, e.g., polytetrafluoroethylene (PTFE).

 After hydrothermal treatment, insoluble material is recovered from the treating vessel. The insoluble material may be recovered by any conventional method, e.g., centrifugation or filtration. Preferably the hydrothermally treated material is cooled

prior to recovery of the insoluble material. If desired, the insoluble material may be washed one or more times with water.

The recovered insoluble material may then be dried by any suitable method known in the art. Such methods include, without limitation, vacuum drying, freeze
5 drying and air drying.

For example, in order to remove any residual water: Vacuum drying is generally performed at pressures ranging from 10mmHg to 500mmHg, with or without the application of heat. Freeze drying typically entails freezing the recovered material, using, for instance, liquid nitrogen, and drying the frozen material under
10 vacuum. Air drying may be effected at temperatures ranging from 25°C to 90°C.

Calcination of the recovered insoluble material may be conducted in an oxidizing atmosphere, e.g., in air, oxygen-enriched air or oxygen, or in the substantial absence of oxygen, e.g., in an inert atmosphere or in vacuo. The inert atmosphere may be any material which is substantially inert, i.e., does not react or interact with, the
15 catalyst precursor. Suitable examples include, without limitation, nitrogen, argon, xenon, helium or mixtures thereof. Preferably, the inert atmosphere is argon or nitrogen. The inert atmosphere may flow over the surface of the catalyst precursor or may not flow thereover (a static environment). When the inert atmosphere does flow over the surface of the catalyst precursor, the flow rate can vary over a wide range,
20 e.g., at a space velocity of from 1 to 500 hr⁻¹.

The calcination is usually performed at a temperature of from 350°C to 850°C, preferably from 400°C to 700°C, more preferably from 500°C to 650°C. The calcination is performed for an amount of time suitable to form the aforementioned catalyst. Typically, the calcination is performed for from 0.5 to 30 hours, preferably
25 from 1 to 25 hours, more preferably for from 1 to 15 hours, to obtain the desired promoted mixed metal oxide.

In a preferred mode of operation, calcination is conducted in two stages, i.e. the recovered insoluble material is heated to a first temperature in an oxidizing atmosphere and, then, the so-treated recovered insoluble material is heated from the
30 first temperature to a second temperature in a non-oxidizing atmosphere. Typically, in the first stage, the catalyst precursor is calcined in an oxidizing environment (e.g. air) at a temperature of from 200°C to 400°C, preferably from 275°C to 325°C for from 15 minutes to 8 hours, preferably for from 1 to 3 hours. In the second stage, the

material from the first stage is calcined in a non-oxidizing environment (e.g., an inert atmosphere) at a temperature of from 500°C to 750°C, preferably for from 550°C to 650°C, for 15 minutes to 8 hours, preferably for from 1 to 3 hours. Optionally, a reducing gas, such as, for example, ammonia or hydrogen, may be added during the
 5 second stage calcination.

In a particularly preferred mode of operation, in the first stage, the material to be calcined is placed in the desired oxidizing atmosphere at room temperature and then raised to the first stage calcination temperature and held there for the desired first stage calcination time. The atmosphere is then replaced with the desired non-oxidizing
 10 atmosphere for the second stage calcination, the temperature is raised to the desired second stage calcination temperature and held there for the desired second stage calcination time.

In an another preferred mode of operation, the insoluble material recovered from the contact mixture is calcined in a non-oxidizing atmosphere, preferably an
 15 inert atmosphere. In the case where the promoter element, i.e. the element X is a halogen, such halogen may be added to the aforementioned non-oxidizing atmosphere.

Although any type of heating mechanism, e.g., a furnace, may be utilized during the calcination, it is preferred to conduct the calcination under a flow of the designated gaseous environment. Therefore, it is advantageous to conduct the
 20 calcination in a bed with continuous flow of the desired gas(es) through the bed of solid catalyst precursor particles.

The starting materials for the above mixed metal oxide are not particularly limited. A wide range of materials including, for example, oxides, nitrates, halides or oxyhalides, alkoxides, acetylacetonates, and organometallic compounds may be used.
 25 For example, ammonium heptamolybdate may be utilized for the source of molybdenum in the catalyst. However, compounds such as MoO_3 , MoO_2 , MoCl_5 , MoOCl_4 , $\text{Mo}(\text{OC}_2\text{H}_5)_5$, molybdenum acetylacetonate, phosphomolybdic acid and silicomolybdic acid may also be utilized instead of ammonium heptamolybdate. Similarly, vanadyl sulfate (VOSO_4) may be utilized for the source of vanadium in the
 30 catalyst. However, compounds such as V_2O_5 , V_2O_3 , VOCl_3 , VCl_4 , $\text{VO}(\text{OC}_2\text{H}_5)_3$, vanadium acetylacetonate and vanadyl acetylacetonate may also be utilized instead of vanadyl sulfate. TeO_2 may be utilized for the source of tellurium in the catalyst. However, TeCl_4 , $\text{Te}(\text{OC}_2\text{H}_5)_5$ and $\text{Te}(\text{OCH}(\text{CH}_3)_2)_4$ may also be utilized instead of

TeO₂. The niobium source may include ammonium niobium oxalate, Nb₂O₅, NbCl₅, niobic acid or Nb(OC₂H₅)₅ as well as the more conventional niobium oxalate.

With respect to the other elements or compounds thereof that may be utilized to make the improved catalysts of the present invention, no particular restrictions are placed thereon. With this in mind, the following lists are merely illustrative of
5 available sources of some of these elements or the compounds thereof, and are not meant to be limiting hereon.

The gold source may be ammonium tetrachloroaurate, gold bromide, gold chloride, gold cyanide, gold hydroxide, gold iodide, gold oxide, gold trichloride acid
10 or gold sulfide.

The silver source may be silver acetate, silver acetylacetonate, silver benzoate, silver bromide, silver carbonate, silver chloride, silver citrate hydrate, silver fluoride, silver iodide, silver lactate, silver nitrate, silver nitrite, silver oxide, silver phosphate, or a solution of silver in an aqueous inorganic acid, e.g., nitric acid.

The copper source may be copper acetate, copper acetate monohydrate, copper acetate hydrate, copper acetylacetonate, copper bromide, copper carbonate, copper chloride, copper chloride dihydrate, copper fluoride, copper formate hydrate, copper gluconate, copper hydroxide, copper iodide, copper methoxide, copper nitrate hydrate, copper nitrate, copper oxide, copper tartrate hydrate, or a solution of copper in an
15 aqueous inorganic acid, e.g., nitric acid.
20

The yttrium source may be an yttrium salt, e.g., yttrium nitrate, dissolved in water.

The scandium source may be scandium acetate, scandium bromide hydrate, scandium chloride, scandium chloride hexahydrate, scandium chloride hydrate,
25 scandium fluoride, scandium iodide, scandium isopropoxide, scandium nitrate hydrate, scandium oxalate hydrate, scandium oxide, or a solution of scandium in an aqueous inorganic acid, e.g., nitric acid.

The rhenium source may be ammonium perrhenate, rhenium carbonyl, rhenium chloride, rhenium fluoride, rhenium oxide, rhenium pentacarbonyl bromide,
30 rhenium pentacarbonyl chloride and rhenium sulfide.

The iridium source may be iridium acetylacetonate, iridium bromide hydrate, iridium chloride, iridium chloride hydrochloride hydrate, iridium chloride hydrate,

iridium oxide, iridium oxide hydrate, iridium oxoacetate trihydrate or iridium dissolved in an aqueous inorganic acid, e.g., nitric acid.

The zinc source may be zinc acetate, zinc acetylacetonate, zinc acrylate, zinc bis(2,2,6,6-tetramethyl-3,5-heptanedioate), zinc bromide, zinc carbonate hydroxide, zinc chloride, zinc citrate, zinc cyclohexanecarboxylate, zinc 3,5-di-tert-butylsalicylate, zinc fluoride, zinc iodide, zinc L-lactate, zinc methacrylate, zinc nitrate, zinc oxide, zinc perchlorate or zinc stearate.

The gallium source may be Ga_2O_3 , GaCl_3 , GaCl_2 , gallium acetylacetonate, Ga_2O_3 or $\text{Ga}(\text{NO}_3)_3$.

The bromine source may be added as one of the above reagents as a bromide, e.g., as the bromide salt of X (where X is an element selected from the group consisting of Li, Na, K, Rb, Cs, Mg, Ca, Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ta, Cr, W, Mn, Re, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, B, Ga, In, Pb, P, As, Sb, Bi, Se, Pr, Nd, Sm and Tb), as molybdenum bromide, as tellurium tetrabromide or as vanadium bromide; or as a solution of bromine in an aqueous inorganic acid, e.g., nitric acid. The bromine source may also be added during the calcination of the recovered insoluble material or, after calcination, as a bromine treatment step. For example, the bromine source may be added to the calcination atmosphere or to the oxidation or ammoxidation reactor feed stream, as, for example, HBr , Br_2 , ethyl bromide or the like, to achieve a promotional effect with the bromine.

The chlorine source may also be added as one of the above reagents as a chloride, e.g., as the chloride salt of X (where X is an element selected from the group consisting of Li, Na, K, Rb, Cs, Mg, Ca, Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ta, Cr, W, Mn, Re, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, B, Ga, In, Pb, P, As, Sb, Bi, Se, Pr, Nd, Sm and Tb), as molybdenum chloride, as tellurium tetrachloride or as vanadium chloride. The chlorine source may also be added during the calcination of the recovered insoluble material or, after calcination, as a chlorine treatment step. For example, the chlorine source may be added to the calcination atmosphere or to the oxidation or ammoxidation reactor feed stream, as, for example, HCl , Cl_2 , ethyl chloride or the like, to achieve a promotional effect with the chloride.

The fluorine source may be added as one of the above reagents as a fluoride, e.g., as the fluoride salt of X (where X is an element selected from the group consisting of Li, Na, K, Rb, Cs, Mg, Ca, Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ta, Cr, W, Mn,

Re, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, B, Ga, In Pb, P, As, Sb, Bi, Se, Pr, Nd, Sm and Tb), as molybdenum fluoride, as tellurium fluoride or as vanadium fluoride. The fluorine source may also be added during the calcination of the recovered insoluble material or, after calcination, as a fluorine treatment step. For example, the fluorine source may be added to the calcination atmosphere or to the oxidation or ammoxidation reactor feed stream, as, for example, HF, F₂, ethyl fluoride or the like, to achieve a promotional effect with the fluoride.

The iodine source may be added as one of the above reagents as an iodide, e.g., as the iodide salt of X (where X is an element selected from the group consisting of Li, Na, K, Rb, Cs, Mg, Ca, Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ta, Cr, W, Mn, Re, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, B, Ga, In Pb, P, As, Sb, Bi, Se, Pr, Nd, Sm and Tb), as molybdenum iodide, as tellurium iodide or as vanadium iodide. The iodine source may also be added during the calcination of the recovered insoluble material or, after calcination, as an iodine treatment step. For example, the iodine source may be added to the calcination atmosphere or to the oxidation or ammoxidation reactor feed stream, as, for example, HI, I₂, ethyl iodide or the like, to achieve a promotional effect with the iodide.

The catalyst formed by the aforementioned techniques, exhibits good catalytic activities by itself. However, the mixed metal oxide may be converted to a catalyst having higher activities by grinding.

There is no particular restriction as to the grinding method, and conventional methods may be employed. As a dry grinding method, a method of using a gas stream grinder may, for example, be mentioned wherein coarse particles are permitted to collide with one another in a high speed gas stream for grinding. The grinding may be conducted not only mechanically but also by using a mortar or the like in the case of a small scale operation.

As a wet grinding method wherein grinding is conducted in a wet state by adding water or an organic solvent to the above mixed metal oxide, a conventional method of using a rotary cylinder-type medium mill or a medium-stirring type mill, may be mentioned. The rotary cylinder-type medium mill is a wet mill of the type wherein a container for the object to be ground is rotated, and it includes, for example, a ball mill and a rod mill. The medium-stirring type mill is a wet mill of the type wherein the object to be ground, contained in a container is stirred by a stirring

apparatus, and it includes, for example, a rotary screw type mill, and a rotary disc type mill.

The conditions for grinding may suitably be set to meet the nature of the above-mentioned promoted mixed metal oxide, the viscosity, the concentration, etc. of the solvent used in the case of wet grinding, or the optimum conditions of the grinding apparatus.

Further, in some cases, it is possible to further improve the catalytic activities by further adding a solvent to the ground catalyst precursor to form a solution or slurry, followed by drying again. There is no particular restriction as to the concentration of the solution or slurry, and it is usual to adjust the solution or slurry so that the total amount of the starting material compounds for the ground catalyst precursor is from 10 to 60 wt %. Then, this solution or slurry is dried by a method such as spray drying, freeze drying, evaporation to dryness or vacuum drying, preferably by the spray drying method. Further, similar drying may be conducted also in the case where wet grinding is conducted.

The oxide obtained by the above-mentioned method may be used as a final catalyst, but it may further be subjected to heat treatment usually at a temperature of from 200° to 700°C for from 0.1 to 10 hours.

The mixed metal oxide thus obtained may be used by itself as a solid catalyst, but may be formed into a catalyst together with a suitable carrier such as silica, alumina, titania, aluminosilicate, diatomaceous earth or zirconia. Further, it may be molded into a suitable shape and particle size depending upon the scale or system of the reactor.

In its fifth and seventh aspects, the present invention provides processes for producing an unsaturated carboxylic acid, which comprise subjecting an alkane, or a mixture of an alkane and an alkene, to a vapor phase catalytic oxidation reaction in the presence of a catalyst according to the first or third aspects of the invention, respectively.

In the production of such an unsaturated carboxylic acid, it is preferred to employ a starting material gas which contains steam. In such a case, as a starting material gas to be supplied to the reaction system, a gas mixture comprising a steam-containing alkane, or a steam-containing mixture of alkane and alkene, and an oxygen-containing gas, is usually used. However, the steam-containing alkane, or the

steam-containing mixture of alkane and alkene, and the oxygen-containing gas may be alternately supplied to the reaction system. The steam to be employed may be present in the form of steam gas in the reaction system, and the manner of its introduction is not particularly limited.

5 In regard to the use of a halogen as an element of the catalyst of the present invention, it is possible to add a gaseous halogen source as previously identified to the gas feed to the reaction.

 Further, as a diluting gas, an inert gas such as nitrogen, argon or helium may be supplied. The molar ratio (alkane or mixture of alkane and alkene) : (oxygen) :
10 (diluting gas) : (H₂O) in the starting material gas is preferably (1) : (0.1 to 10) : (0 to 20) : (0.2 to 70), more preferably (1) : (1 to 5.0) : (0 to 10) : (5 to 40).

 When steam is supplied together with the alkane, or the mixture of alkane and alkene, as starting material gas, the selectivity for an unsaturated carboxylic acid is distinctly improved, and the unsaturated carboxylic acid can be obtained from the
15 alkane, or mixture of alkane and alkene, in good yield simply by contacting in one stage. However, the conventional technique utilizes a diluting gas such as nitrogen, argon or helium for the purpose of diluting the starting material. As such a diluting gas, to adjust the space velocity, the oxygen partial pressure and the steam partial pressure, an inert gas such as nitrogen, argon or helium may be used together with the
20 steam.

 As the starting material alkane it is preferred to employ a C₃₋₈alkane, particularly propane, isobutane or n-butane; more preferably, propane or isobutane; most preferably, propane. According to the present invention, from such an alkane, an unsaturated carboxylic acid such as an α,β -unsaturated carboxylic acid can be
25 obtained in good yield. For example, when propane or isobutane is used as the starting material alkane, acrylic acid or methacrylic acid will be obtained, respectively, in good yield.

 In the present invention, as the starting material mixture of alkane and alkene, it is possible to employ a mixture of C₃₋₈alkane and C₃₋₈alkene, particularly propane
30 and propene, isobutane and isobutene or n-butane and n-butene. As the starting material mixture of alkane and alkene, propane and propene or isobutane and isobutene are more preferred. Most preferred is a mixture of propane and propene. According to the present invention, from such a mixture of an alkane and an alkene,

an unsaturated carboxylic acid such as an α,β -unsaturated carboxylic acid can be obtained in good yield. For example, when propane and propene or isobutane and isobutene are used as the starting material mixture of alkane and alkene, acrylic acid or methacrylic acid will be obtained, respectively, in good yield. Preferably, in the mixture of alkane and alkene, the alkene is present in an amount of at least 0.5% by weight, more preferably at least 1.0% by weight to 95% by weight; most preferably, 3% by weight to 90% by weight.

As an alternative, an alkanol, such as isobutanol, which will dehydrate under the reaction conditions to form its corresponding alkene, i.e. isobutene, may also be used as a feed to the present process or in conjunction with the previously mentioned feed streams.

The purity of the starting material alkane is not particularly limited, and an alkane containing a lower alkane such as methane or ethane, air or carbon dioxide, as impurities, may be used without any particular problem. Further, the starting material alkane may be a mixture of various alkanes. Similarly, the purity of the starting material mixture of alkane and alkene is not particularly limited, and a mixture of alkane and alkene containing a lower alkene such as ethene, a lower alkane such as methane or ethane, air or carbon dioxide, as impurities, may be used without any particular problem. Further, the starting material mixture of alkane and alkene may be a mixture of various alkanes and alkenes.

There is no limitation on the source of the alkene. It may be purchased, per se, or in admixture with an alkane and/or other impurities. Alternatively, it can be obtained as a by-product of alkane oxidation. Similarly, there is no limitation on the source of the alkane. It may be purchased, per se, or in admixture with an alkene and/or other impurities. Moreover, the alkane, regardless of source, and the alkene, regardless of source, may be blended as desired.

The detailed mechanism of the oxidation reaction of the present invention is not clearly understood, but the oxidation reaction is carried out by oxygen atoms present in the above promoted mixed metal oxide or by molecular oxygen present in the feed gas. To incorporate molecular oxygen into the feed gas, such molecular oxygen may be pure oxygen gas. However, it is usually more economical to use an oxygen-containing gas such as air, since purity is not particularly required.

It is also possible to use only an alkane, or a mixture of alkane and alkene, substantially in the absence of molecular oxygen for the vapor phase catalytic reaction. In such a case, it is preferred to adopt a method wherein a part of the catalyst is appropriately withdrawn from the reaction zone from time to time, then sent to an oxidation regenerator, regenerated and then returned to the reaction zone for reuse. As the regeneration method of the catalyst, a method may, for example, be mentioned which comprises contacting an oxidative gas such as oxygen, air or nitrogen monoxide with the catalyst in the regenerator usually at a temperature of from 300° to 600°C.

This aspect of the present invention will be described in further detail with respect to a case where propane is used as the starting material alkane and air is used as the oxygen source. The reaction system may be a fixed bed system or a fluidized bed system. However, since the reaction is an exothermic reaction, a fluidized bed system may preferably be employed whereby it is easy to control the reaction temperature. The proportion of air to be supplied to the reaction system is important for the selectivity for the resulting acrylic acid, and it is usually at most 25 moles, preferably from 0.2 to 18 moles per mole of propane, whereby high selectivity for acrylic acid can be obtained. This reaction can be conducted usually under atmospheric pressure, but may be conducted under a slightly elevated pressure or slightly reduced pressure. With respect to other alkanes such as isobutane, or to mixtures of alkanes and alkenes such as propane and propene, the composition of the feed gas may be selected in accordance with the conditions for propane.

Typical reaction conditions for the oxidation of propane or isobutane to acrylic acid or methacrylic acid may be utilized in the practice of the present invention. The process may be practiced in a single pass mode (only fresh feed is fed to the reactor) or in a recycle mode (at least a portion of the reactor effluent is returned to the reactor). General conditions for the process of the present invention are as follows: the reaction temperature can vary from 200°C to 700°C, but is usually in the range of from 200°C to 550°C, more preferably 250°C to 480°C, most preferably 300°C to 400°C; the gas space velocity, SV, in the vapor phase reaction is usually within a range of from 100 to 10,000 hr⁻¹, preferably 300 to 6,000 hr⁻¹, more preferably 300 to 2,000 hr⁻¹; the average contact time with the catalyst can be from 0.01 to 10 seconds

or more, but is usually in the range of from 0.1 to 10 seconds, preferably from 2 to 6 seconds; the pressure in the reaction zone usually ranges from 0 to 75 psig, but is preferably no more than 50 psig. In a single pass mode process, it is preferred that the oxygen be supplied from an oxygen-containing gas such as air. The single pass mode process may also be practiced with oxygen addition. In the practice of the recycle mode process, oxygen gas by itself is the preferred source so as to avoid the build up of inert gases in the reaction zone.

Of course, in the oxidation reaction of the present invention, it is important that the hydrocarbon and oxygen concentrations in the feed gases be maintained at the appropriate levels to minimize or avoid entering a flammable regime within the reaction zone or especially at the outlet of the reactor zone. Generally, it is preferred that the outlet oxygen levels be low to both minimize after-burning and, particularly, in the recycle mode of operation, to minimize the amount of oxygen in the recycled gaseous effluent stream. In addition, operation of the reaction at a low temperature (below 450°C) is extremely attractive because after-burning becomes less of a problem which enables the attainment of higher selectivity to the desired products. The catalyst of the present invention operates more efficiently at the lower temperature range set forth above, significantly reducing the formation of acetic acid and carbon oxides, and increasing selectivity to acrylic acid. As a diluting gas to adjust the space velocity and the oxygen partial pressure, an inert gas such as nitrogen, argon or helium may be employed.

When the oxidation reaction of propane, and especially the oxidation reaction of propane and propene, is conducted by the method of the present invention, carbon monoxide, carbon dioxide, acetic acid, etc. may be produced as by-products, in addition to acrylic acid. Further, in the method of the present invention, an unsaturated aldehyde may sometimes be formed depending upon the reaction conditions. For example, when propane is present in the starting material mixture, acrolein may be formed; and when isobutane is present in the starting material mixture, methacrolein may be formed. In such a case, such an unsaturated aldehyde can be converted to the desired unsaturated carboxylic acid by subjecting it again to the vapor phase catalytic oxidation with the promoted mixed metal oxide-containing catalyst of the present invention or by subjecting it to a vapor phase catalytic oxidation reaction with a conventional oxidation reaction catalyst for an unsaturated aldehyde.

In its sixth and eighth aspects, the present invention provides processes for producing an unsaturated nitrile, which comprise subjecting an alkane, or a mixture of an alkane and an alkene, to a vapor phase catalytic oxidation reaction with ammonia in the presence of a catalyst according to the first or third aspects of the invention,
5 respectively.

In regard to the use of a halogen as an element of the catalyst of the present invention, it is possible to add a gaseous halogen source as previously identified to the gas feed to the reaction.

In the production of such an unsaturated nitrile, as the starting material alkane,
10 it is preferred to employ a C₃₋₈alkane such as propane, butane, isobutane, pentane, hexane and heptane. However, in view of the industrial application of nitriles to be produced, it is preferred to employ a lower alkane having 3 or 4 carbon atoms, particularly propane and isobutane.

Similarly, as the starting material mixture of alkane and alkene, it is possible
15 to employ a mixture of C₃₋₈alkane and C₃₋₈alkene such as propane and propene, butane and butene, isobutane and isobutene, pentane and pentene, hexane and hexene, and heptane and heptene. However, in view of the industrial application of nitriles to be produced, it is more preferred to employ a mixture of a lower alkane having 3 or 4 carbon atoms and a lower alkene having 3 or 4 carbon atoms, particularly propane and
20 propene or isobutane and isobutene. Preferably, in the mixture of alkane and alkene, the alkene is present in an amount of at least 0.5% by weight, more preferably at least 1.0% by weight to 95% by weight, most preferably 3% by weight to 90% by weight.

The purity of the starting material alkane is not particularly limited, and an alkane containing a lower alkane such as methane or ethane, air or carbon dioxide, as
25 impurities, may be used without any particular problem. Further, the starting material alkane may be a mixture of various alkanes. Similarly, the purity of the starting material mixture of alkane and alkene is not particularly limited, and a mixture of alkane and alkene containing a lower alkene such as ethene, a lower alkane such as methane or ethane, air or carbon dioxide, as impurities, may be used without any
30 particular problem. Further, the starting material mixture of alkane and alkene may be a mixture of various alkanes and alkenes.

There is no limitation on the source of the alkene. It may be purchased, per se, or in admixture with an alkane and/or other impurities. Alternatively, it can be

obtained as a by-product of alkane oxidation. Similarly, there is no limitation on the source of the alkane. It may be purchased, per se, or in admixture with an alkene and/or other impurities. Moreover, the alkane, regardless of source, and the alkene, regardless of source, may be blended as desired.

5 The detailed mechanism of the ammoxidation reaction of this aspect of the present invention is not clearly understood. However, the oxidation reaction is conducted by the oxygen atoms present in the above promoted mixed metal oxide or by the molecular oxygen in the feed gas. When molecular oxygen is incorporated in the feed gas, the oxygen may be pure oxygen gas. However, since high purity is not
10 required, it is usually economical to use an oxygen-containing gas such as air.

As the feed gas, it is possible to use a gas mixture comprising an alkane, or a mixture of an alkane and an alkene, ammonia and an oxygen-containing gas, However, a gas mixture comprising an alkane or a mixture of an alkane and an alkene and ammonia, and an oxygen-containing gas may be supplied alternately.

15 When the gas phase catalytic reaction is conducted using an alkane, or a mixture of an alkane and an alkene, and ammonia substantially free from molecular oxygen, as the feed gas, it is advisable to employ a method wherein a part of the catalyst is periodically withdrawn and sent to an oxidation regenerator for regeneration, and the regenerated catalyst is returned to the reaction zone. As a
20 method for regenerating the catalyst, a method may be mentioned wherein an oxidizing gas such as oxygen, air or nitrogen monoxide is permitted to flow through the catalyst in the regenerator usually at a temperature of from 300°C to 600°C.

This aspect of the present invention will be described in further detail with respect to a case where propane is used as the starting material alkane and air is used
25 as the oxygen source. The proportion of air to be supplied for the reaction is important with respect to the selectivity for the resulting acrylonitrile. Namely, high selectivity for acrylonitrile is obtained when air is supplied within a range of at most 25 moles, particularly 1 to 15 moles, per mole of the propane. The proportion of ammonia to be supplied for the reaction is preferably within a range of from 0.2 to 5 moles,
30 particularly from 0.5 to 3 moles, per mole of propane. This reaction may usually be conducted under atmospheric pressure, but may be conducted under a slightly increased pressure or a slightly reduced pressure. With respect to other alkanes such as isobutane, or to mixtures of alkanes and alkenes such as propane and propene, the

composition of the feed gas may be selected in accordance with the conditions for propane.

The process of this aspect of the present invention may be conducted at a temperature of, for example, from 250°C to 480°C. More preferably, the temperature is from 300°C to 400°C. The gas space velocity, SV, in the gas phase reaction is usually within the range of from 100 to 10,000 hr⁻¹, preferably from 300 to 6,000 hr⁻¹, more preferably from 300 to 2,000 hr⁻¹. As a diluent gas, for adjusting the space velocity and the oxygen partial pressure, an inert gas such as nitrogen, argon or helium can be employed. When ammoxidation of propane is conducted by the method of the present invention, in addition to acrylonitrile, carbon monoxide, carbon dioxide, acetonitrile, hydrocyanic acid and acrolein may form as by-products.

Examples

Comparative Example 1

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of Mo_{1.0}V_{0.43}Te_{0.33}Nb_{0.1}O_x. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Comparative Example 2

200 mL of an aqueous solution containing ammonium heptamolybdate tetrahydrate (1.0M Mo), ammonium metavanadate (0.3M V) and telluric acid (0.23M Te) formed by dissolving the corresponding salts in water at 70°C, was added to a 1000mL rotavap flask. Then 100 mL of an aqueous solution of ammonium niobium oxalate

(0.25M Nb) and oxalic acid (0.31M) were added thereto. After removing the water via a rotary evaporator with a warm water bath at 50°C and 28mm Hg, the solid materials were further dried in a vacuum oven at 25°C overnight, and calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{Mo}_{1.0}\text{V}_{0.3}\text{Te}_{0.23}\text{Nb}_{0.125}\text{O}_x$. 20g of the catalyst was ground and added to 400mL of 5 % aqueous HNO_3 . The resulting suspension was stirred at 25°C for 30 minutes, then the solids were collected by gravity filtration, dried in a vacuum oven overnight at 25°C, and heated in argon from 25°C to 600°C at 10°C/min and held at 600°C for two hours. The catalyst, thus obtained, was pressed and sieved to 14 - 20 mesh granules for reactor evaluation. Figure 1 shows the x-ray diffraction pattern of this catalyst.

Comparative Example 3

Ammonium heptamolybdate (18.38 g), ammonium metavanadate (3.65g) and telluric acid (5.50 g) were dissolved in warm water (307g). A solution of 4.96g of ammonium niobium oxalate in water (70.5g) and 9.2g of oxalic acid were added to the first solution. After removing the water via a rotary evaporator with a warm water bath at 50°C and 28mm Hg, the solid materials were further dried in a vacuum oven at 25°C overnight, and calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{Mo}_{1.0}\text{V}_{0.3}\text{Te}_{0.23}\text{Nb}_{0.08}\text{O}_x$. The catalyst, thus obtained, was pressed and sieved to 14 - 20 mesh granules for reactor evaluation. Figure 2 shows the x-ray diffraction pattern of this catalyst.

Example 1

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M palladium nitrate hydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were

hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{Pd}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation. Figure 3 shows the x-ray diffraction pattern of this catalyst.

10 **Example 2**

20g of the calcined solid from Example 1, 20g of oxalic acid dihydrate and 80g of water were added to a 300cc autoclave. The mixture was heated with stirring for 5 hours at 125°C. The solids were collected by gravity filtration, washed with water and dried under vacuum overnight at 25°C. The catalyst thus obtained was pressed and
15 sieved to 14-20 mesh granules for reactor evaluation. Figure 4 shows the x-ray diffraction pattern of this catalyst.

Example 3

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in
20 water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M tetraamineplatinum(II) nitrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb
25 contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of
30 $\text{Pt}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 4

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M rhodium nitrate dihydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{Rh}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

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Example 5

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M lithium nitrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{Li}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

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Example 6

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs,

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and then 6 mL of 0.1M potassium nitrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $K_{0.01}Mo_{1.0}V_{0.43}Te_{0.33}Nb_{0.1}O_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 7

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M magnesium nitrate hexahydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $Mg_{0.01}Mo_{1.0}V_{0.43}Te_{0.33}Nb_{0.1}O_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 8

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M calcium nitrate tetrahydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the

bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of

5 $\text{Ca}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 9

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in

10 water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M samarium nitrate hexahydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb

15 contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of

20 $\text{Sm}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 10

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in

25 water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M chromium nitrate nonahydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb

30 contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of

$\text{Cr}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 11

5 To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M manganese (II) nitrate hydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous
10 solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min
15 and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{Mn}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 12

20 To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M cobalt (II) nitrate hexahydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous
25 solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min
30 and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{Co}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 13

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M nickel (II) nitrate hexahydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{Ni}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

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Example 14

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M indium (III) nitrate pentahydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{In}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

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Example 15

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs,

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and then 6 mL of 0.1M lanthanum nitrate hexahydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{La}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 16

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M zinc nitrate hexahydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{Zn}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 17

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M iron (III) nitrate nonahydrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the

bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of

5 $\text{Fe}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 18

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of

10 tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of 0.1M lead (II) nitrate in water and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were

15 hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of

20 $\text{Pb}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 19

To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of

25 tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of vanadium (III) iodide in water (0.1M in I) and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb

30 contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of

$\text{I}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 20

5 To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of molybdenum (III) bromide in water (0.1M in Br) and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of
10 aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at
15 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{Br}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Example 21

20 To a 125 mL Parr Acid Digestion Bomb with an inner tube made of PTFE, 3.15g of tellurium dioxide and 60 mL of 0.143M ammonium heptamolybdate tetrahydrate in water were added. The mixture was first hydrothermally treated at 100°C for 1.5hrs, and then 6 mL of molybdenum (V) chloride in water (0.1M in Cl) and 6.5g of vanadyl sulfate hydrate were added to the bomb at 60°C followed by 30 mL of
25 aqueous solution (0.2M in Nb) of ammonium niobium oxalate with stirring. The bomb contents were hydrothermally treated at 175°C for 4 days. Black solids formed in the bomb were collected by gravity filtration, washed with deionized water (50 mL), dried in a vacuum oven at 25°C overnight, and then calcined in air from 25 to 275°C at 10°C/min and held at 275°C for 1 hour, then in argon from 275 to 600°C at
30 2°C/min and held at 600°C for 2 hours. The final catalyst had a nominal composition of $\text{Cl}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$. The catalyst thus obtained was pressed and sieved to 14-20 mesh granules for reactor evaluation.

Evaluation procedure

Catalysts were evaluated in a 10 cm long Pyrex tube reactor (internal diameter: 3.9 mm). The catalyst bed (4 cm long) was positioned with glass wool at approximately mid-length in the reactor and was heated with an electric furnace. Mass flow controllers and meters regulated the gas flow rate. The oxidation was conducted using a feed gas stream of propane, steam and air, with a feed ratio of propane:steam:air of 1:3:96. The reactor effluent was analyzed by an FTIR. The results at a 3 second residence time are shown in Tables 1-2.

Table 1

Comp. Example	Composition	Temp. (°C)	% C3 Conv.	% AA Yield
1	$\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	393	68	34
2	$\text{Mo}_{1.0}\text{V}_{0.3}\text{Te}_{0.23}\text{Nb}_{0.125}\text{O}_x$	384	75	52
3	$\text{Mo}_{1.0}\text{V}_{0.3}\text{Te}_{0.23}\text{Nb}_{0.08}\text{O}_x$	400	6	2

Table 2

Example	Composition	Temp. (°C)	% C3 Conv.	% AA Yield
1	$\text{Pd}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	390	70	48
2	$\text{Pd}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	386	80	50
3	$\text{Pt}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	376	54	43
4	$\text{Rh}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	341	46	41
5	$\text{Li}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	375	57	44
6	$\text{K}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	327	52	44
7	$\text{Mg}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	391	55	44
8	$\text{Ca}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	341	47	41
9	$\text{Sm}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	369	60	43
10	$\text{Cr}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	380	56	40
11	$\text{Mn}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	383	59	45
12	$\text{Co}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	350	65	45
13	$\text{Ni}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	372	52	42
14	$\text{In}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	374	42	39
15	$\text{La}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	344	60	39
16	$\text{Zn}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	369	46	38
17	$\text{Fe}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	375	47	38
18	$\text{Pb}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	336	44	37
19	$\text{I}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	400	39	32
20	$\text{Br}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	379	49	35
21	$\text{Cl}_{0.01}\text{Mo}_{1.0}\text{V}_{0.43}\text{Te}_{0.33}\text{Nb}_{0.1}\text{O}_x$	370	50	25